

# Material Science Callister 9th Edition Solution

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42 For some hypothetical metal, the equilibrium number of vacancies at 900 C is  $23 \times 10^{25} \text{ m}^{-3}$ . If the density and atomic weight of this metal are  $740 \text{ g/cm}^3$  and  $855 \text{ g/mol}$ , respectively, calculate the fraction of vacancies for this metal at 900 C. Solution This problem is solved using two steps: (1) calculate the total number of lattice sites in silver,  $N$ , using

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25 (a) How many grams are there in one amu of a material? (b) Mole, in the context of this book, is taken in units of gram-mole. On this basis, how many atoms are there in a pound-mole of a substance? Solution (a) In order to determine the number of grams in one amu of material, appropriate manipulation of the

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22 Silicon has three naturally occurring isotopes: 9223% of  $^{28}\text{Si}$ , with an atomic weight of 279769 amu, 468% of  $^{29}\text{Si}$ , with an atomic weight of 289765 amu, and 309% of  $^{30}\text{Si}$ , with an atomic weight of 299738 amu On the basis of these data, confirm that the average atomic weight of Si is 280854 amu

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35 Show that the atomic packing factor for BCC is 068 Solution The atomic packing factor is defined as the ratio of sphere volume ( $V_S$ ) to the total

unit cell volume (VC ), or Because there are two spheres associated with each unit cell for BCC

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